

Spring

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STATE OF ALABAMA
DIVISION OF THE ATTORNEY GENERAL
OFFICE OF THE ATTORNEY GENERAL
MONTGOMERY, ALABAMA

IN RE: [Illegible Name]
[Illegible Address]
[Illegible City, State, Zip]

Case No. [Illegible]

FILED FOR RECORD IN THE [Illegible] COUNTY COURT AT [Illegible] ALABAMA

THIS [Illegible] DAY OF [Illegible] 20[Illegible]

ATTEST: [Illegible Signature]

[Illegible Name], [Illegible Title]

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THE RICORAST STOCK COLLECTION OF NEWSPAPER

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THE UNIVERSITY OF CALIFORNIA, BERKELEY
 DEPARTMENT OF CHEMISTRY
 CHEMISTRY 104B
 LABORATORY REPORT

NAME: _____
 STUDENT ID: _____
 DATE: _____

1. **INTRODUCTION**
 The purpose of this experiment is to determine the molar mass of a volatile liquid by measuring its mass and volume under known conditions of temperature and pressure. This is done by vaporizing a known mass of the liquid in a flask of known volume and measuring the mass of the condensed liquid.

2. **PROCEDURE**
 A small amount of the volatile liquid was placed in a flask of known volume. The flask was heated in a boiling water bath until the liquid had completely vaporized and the flask was filled with the vapor. The flask was then cooled, and the condensed liquid was weighed. The mass of the condensed liquid was compared to the mass of the empty flask to determine the mass of the liquid.

3. **RESULTS**
 The mass of the condensed liquid was found to be 0.456 g. The volume of the flask was 125 mL. The temperature of the boiling water bath was 100°C. The pressure was 1 atm.

4. **DISCUSSION**
 The molar mass of the liquid can be calculated from the ideal gas law. The mass of the liquid is 0.456 g, the volume is 0.125 L, the temperature is 373 K, and the pressure is 1 atm. The molar mass is calculated to be 104.1 g/mol.

5. **CONCLUSION**
 The molar mass of the volatile liquid is 104.1 g/mol. This value is consistent with the molar mass of acetone, which is 58.08 g/mol. The discrepancy is likely due to experimental error.

6. **REFERENCES**
 No references were used in this experiment.

7. **APPENDIX**
 The following table shows the data for this experiment.

Mass of condensed liquid (g)	Volume of flask (L)	Temperature (K)	Pressure (atm)	Molar mass (g/mol)
0.456	0.125	373	1	104.1